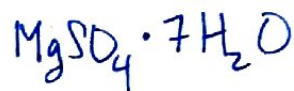


# Unit 1 Test Review Kahoot

#1.)  $250 - 120 = 130 \text{ g H}_2\text{O} \times \frac{1 \text{ mol}}{18 \text{ g}} \approx 7.2 \text{ mol}$

$120 \text{ g} \times \frac{1 \text{ mol}}{120 \text{ g}} = 1 \text{ mol MgSO}_4$



#3.)  $1.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol}}{18 \text{ g}} \times \frac{2 \text{ H}}{1 \text{ H}_2\text{O}} = 0.2 \text{ mol H}$

$8.8 \text{ g CO}_2 \times \frac{1 \text{ mol}}{44 \text{ g}} \times \frac{1 \text{ C}}{1 \text{ CO}_2} = 0.2 \text{ mol C}$

= 1

÷ 0.2

= 1



#5.)  $M = \frac{\text{mol}}{\text{L}}$

$1.2 \text{ M} = \frac{\text{mol}}{0.52 \text{ L}}$

$\Rightarrow (1.2)(0.52) = 0.62 \text{ mol Ba(OH)}_2$

$0.62 \text{ mol Ba(OH)}_2$

$\times \frac{3 \text{ mol ions}}{1 \text{ mol Ba(OH)}_2}$

$= 1.9 \text{ mol ions}$

#6.)  $96 \text{ g C} \times \frac{1 \text{ mol}}{12 \text{ g}} = 8 \text{ mol C}$

$8 \text{ g H} \times \frac{1 \text{ mol}}{1 \text{ g}} = 8 \text{ mol H}$

÷ 8

= 1

= 1



$\frac{78}{13} = 6$