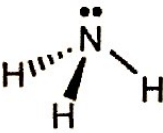
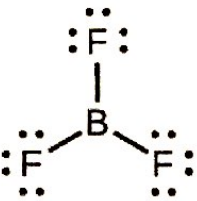
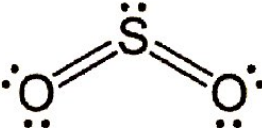
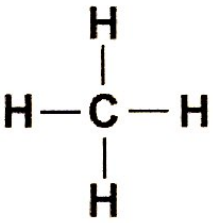
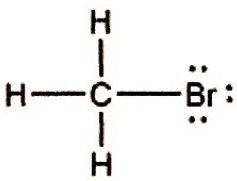
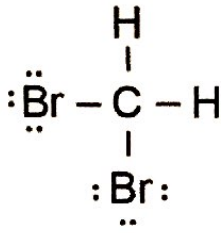
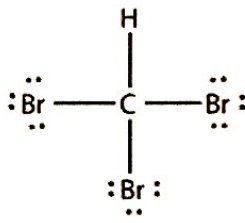
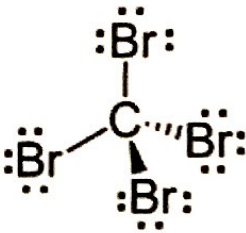


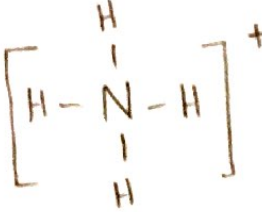
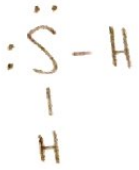
Let's Practice!

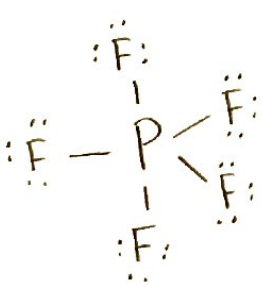
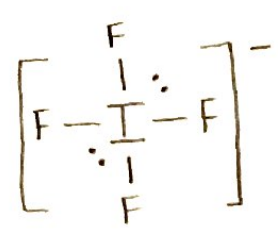
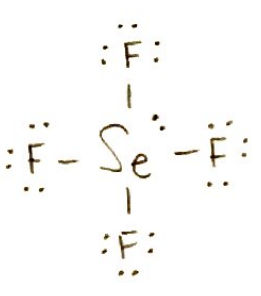
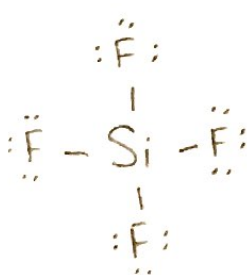
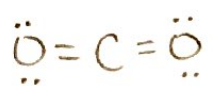
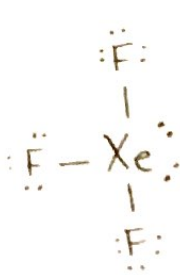
Given the Lewis dot structures below, determine if the molecule is polar or non-polar.

1)  Is the molecule: <u>polar</u> or non-polar?	2)  Is the molecule: polar or <u>non-polar</u> ?	3)  Is the molecule: <u>polar</u> or non-polar?	4)  Is the molecule: polar or <u>non-polar</u> ?
5)  Is the molecule: <u>polar</u> or non-polar?	6)  Is the molecule: <u>polar</u> or non-polar?	7)  Is the molecule: <u>polar</u> or non-polar?	8)  Is the molecule: polar or <u>non-polar</u> ?

And... Even More Practice!

For each of the compounds given below, identify its VSEPR shape, bond angle(s), the hybridization of its central atom, and its molecular polarity.

CO	NH ₄ ⁺	H ₂ S
Lewis structure: $:C \equiv O:$	Lewis structure: 	Lewis structure: 
Molecular geometry: <u>linear</u> Bond angle(s): <u>180°</u> Central Atom Hybridization: <u>sp</u> Polar or not? <u>polar</u>	Molecular geometry: <u>tetrahedral</u> Bond angle(s): <u>109°</u> Central Atom Hybridization: <u>sp³</u> Polar or not? <u>non-polar</u>	Molecular geometry: <u>bent</u> Bond angle(s): <u>~105°</u> Central Atom Hybridization: <u>sp³</u> Polar or not? <u>polar</u>

PF ₅	IF ₄ ⁻	SeF ₄
Lewis structure: 	Lewis structure: 	Lewis structure: 
Molecular geometry: <u>trigonal bipyramidal</u> Bond angle(s): <u>90°, 120°</u> Central Atom Hybridization: <u>sp³d</u> Polar or not? <u>non-polar</u>	Molecular geometry: <u>square planar</u> Bond angle(s): <u>90°</u> Central Atom Hybridization: <u>sp³d²</u> Polar or not? <u>non-polar</u>	Molecular geometry: <u>See-saw</u> Bond angle(s): <u>~118°, ~88°</u> Central Atom Hybridization: <u>sp³d</u> Polar or not? <u>polar</u>
SiF ₄	CO ₂	XeF ₃ ⁺
Lewis structure: 	Lewis structure: 	Lewis structure: 
Molecular geometry: <u>tetrahedral</u> Bond angle(s): <u>109°</u> Central Atom Hybridization: <u>sp³</u> Polar or not? <u>non-polar</u>	Molecular geometry: <u>linear</u> Bond angle(s): <u>180°</u> Central Atom Hybridization: <u>sp</u> Polar or not? <u>non-polar</u>	Molecular geometry: <u>T-shaped</u> Bond angle(s): <u>90°</u> Central Atom Hybridization: <u>sp³d</u> Polar or not? <u>polar</u>